SSC PUBLIC EXAMINATIONS 2026 : MODEL PAPER -1 : BLUEPRINT

Subject: GENERAL SCIENCE-1 **Class:** 10 TH CLASS

Unit/Paper: **Max Marks:** 50 MARKS **Time:** 2 Hr.15 min.

S. No.	Objective → / Form of Questions / Content Unit / Sub Unit	Knowledge (20%)				Understanding (25%)				Application (20%)				Analysis (15%)				Evaluation (10%)				Creation (10%)				Total (Row-wise)
		E/LA 8m	SA 4m	VSA 2m	O 1m	E/LA 8m	SA 4m	VSA 2m	O 1m	E/LA 8m	SA 4m	VSA 2m	O 1m	E/LA 8m	SA 4m	VSA 2m	O 1m	E/LA 8m	SA 4m	VSA 2m	O 1m	E/LA 8m	SA 4m	VSA 2m	O 1m	
1	Chemical Reactions and Equations													1											1	9
2	Acids Bases and Salts								1		2														1	10
3	Metals and non metals					1																			1	9
4	Carbon and it's compounds			1					1					1												11
5	Light-Reflection and Refraction										1	1							1		1					11
6	The Human Eye and Colourful world	1							1																	9
7	Electricity	1							1															1		11
8	Magnetic effects of electric current					1																				8
		16		2		16			4		12	2		16					4		1			2	3	
	Total (Col-wise) →	10 (8)				12 (8)				10 (4)				8 (8)				5				5				50 (+28 as choice)

Notes: Figures within brackets to indicate the marks of question given as internal choice.

Summary:						
Essay (E)	:	No.	3	**Marks:**	24	
Short Answer (SA)	:	No.	3	**Marks:**	12	
Very Short Answer (VSA)	:	No.	3	**Marks:**	6	
Objective (0)	:	No.	8	**Marks:**	8	

Pattern of Options : Internal Choice for Q.No.s 12, 15,16 and 17

Scheme of Sections 4 (8x1=8; 3x2=6; 3 x4=12; 3x8=24)

Source: BSE Andhra Pradesh (bse.ap.gov.in) • clearsteps.co.in

Format of Design (Subject other than language)

Question Paper/Test

Subject: General science-1
Unit/Paper: 19T/E
Class: X
Time: 2 hrs 15 min
Marks: 50

Weightage to Objective

Objective	Knowledge	Understand	Application	Analyse	Evaluation	Creation	Total
% of Marks	20	25	20	15	10	10	100
Marks	10	12	10	8	5	5	50

Weightage to Form of Question

Forms of Questions	E/LA	SA	VSA	O (MCQ) - 1 Mark Qns	Total
No. of Questions	3	3	3	8	17
Marks Allotted	24	12	6	8	50
Estimated Time (min)	60	30	15	15	120 min

Weightage to Objective

S. No.	Unit/Sub-Units	Marks
1	Chemical Reactions and Equations	9
2	Acids Bases and Salts	10
3	Metals and non metals	9
4	Carbon and it's compounds	11
5	Light-Reflection and Refraction	11
6	The Human Eye and Colourful world	9
7	Electricity	11
8	Magnetic effects of electric current	8
Total		50 + 28

Estimated Difficulty Level	Difficult	15 % Marks
	Average	45 % Marks
	Easy	40 % Marks

Index of Abbreviations

(E/LA: Essay/Long Answer; **SA:** Short Answer; **VSA:** Very Short Answer; **O:** Objective)

SSC PUBLIC EXAMATINATIONS 2025 - 26

GENERAL SCIENCE - PAPER-I

PHYSICAL SCIENCE (MODEL PAPER - 1)

(ENGLISH VERSION)

Time : 2 Hours **Max. Marks : 50**

Instructions :

1. Question paper consists of 4 sections and 17 Questions.
2. Internal Choice is there only for Q.No. 12 in Section-III and for all the Questions in Section-IV.
3. In the duration of 2 hours, 15 minutes of time is allotted to read the Question paper.
4. All answers should be written in the answer booklet only.
5. Answer should be written neatly and legibly.

SECTION-I

Note : 1) Answer all the Questions.
2) Each question carries 1 mark. **(8 × 1 = 8 M)**

1. A magnesium ribbon is burnt in the presence of Oxygen to give Magnesium oxide. Rewrite the above reaction as Chemical equation.
2. Give an example for an acid.
3. The pH value of a solution is 10. What is it's colour in the presence of methyl orange indicator?
4. Propose a method to extract a highly reactive metal from it's ore?
5. Which of the following hydrocarbon undergoes addition reaction?
   A) C_2H_6 B) C_3H_8 C) CH_4 D) C_3H_6
6. The radius of curvature of a spherical mirror is given as 20 cm then Determine it's focal length.
7. The change in focal length of eye lens is caused by the action of the _____ in human eye.
8. 2Ω, 4Ω resistors are connected in series. What will be the resultant resistance ?

SECTION-II

Note : 1) Answer all the Questions.
2) Each question carries 2 marks. **(3 × 2 = 6 M)**

9. What are the two properties of carbon, which lead to the huge number of carbon compounds we see around us?
10. Write any two daily life applications of Lenses.
11. Predict and write why the series arrangement is not used for domestic circuits?

SECTION-III

Note : 1) Answer all the Questions.
2) Each question carries 4 marks. **(3 × 4 = 12 M)**

12 Draw any one of the following diagrams:

(A) Draw a neat ray diagram of formation of image when the object is placed in front of a concave mirror in the following positions :

i) At C ii) between F and C

(OR)

B) Draw a neat diagram to show the reaction of metals with acids.

13. Write the uses of Bleaching powder.

14. Different media and their Refractive indices were given in the following table. Based on the data given, answer the questions given below.

Material Medium	Refractive Index
Air	1.0003
Ice	1.31
Water	1.33
Benzene	1.50

i) Arrange the above material media in the ascending order with respect to the speed of light.

ii) Which material media is optically denser in between Water and Benzene?

iii) Determine the speed of light in Benzene. (C = 3 x 10^8 m/s)

iv) Which is the most optically rarer medium from the table.

SECTION-IV

Note : 1) Answer all the Questions.
2) Each question carries 8 marks.
3) Each question has an internal choice. **(3 × 8 = 24 M)**

15. A) What is Hypermetropia? How do you rectify it ? Write indetailed.

(or)

B) Define the following.

1) Electric current 2) Potential difference

3) Ohm's law 4) Electric power

16. A) Balance the following chemical equations.

1) $NaCl + AgNO_3 \rightarrow AgCl + NaNO_3$

2) $BaCl_2 + H_2SO_4 \rightarrow BaSO_4 + HCl$

3) $Pb(NO_3)_2 \xrightarrow{\Delta} PbO + NO_2 + O_2$ (on Heating)

4) $Fe + H_2O \rightarrow Fe_3O_4 + H_2$

(OR)

Complete the following table.

No. of carbons in Hydro carbon	Alkane	Alkene	Alkyne
3	CH_4		
4		C_4H_8	
5			C_5H_8
6		C_6H_{12}	

17. A) Explain the procedure to show that compass needle is deflected on passing electric current through a metallic conductor. (Oersted's experiment)

(or)

B) Write an activity to show that high reactive metals displaces low reactive metals from their compounds.